Syllabus for Chemistry 101
1:40 – 2:30 pm MWF
Loyola University: Fall 2008

Instructor: Dr. Conrad Naleway; Office FH 422
Office Hours: MW 2:30-3:30pm and by Appointment

Meeting Times; Days & Rooms

Lecture: 1:40-2:30 pm, MWF in FH 133
Discussion & Quizzes:

1086	CHEM	101	002	DIS	9:20AM	10:10AM	Wed	FH	7
1087	CHEM	101	003	DIS	10:25AM	11:15AM	Wed	FH	7
1088	CHEM	101	004	DIS	11:30AM	12:20PM	Wed	FH	7

Review Sessions: There will be weekly optional review sessions and additional review sessions prior to each exam. ***Dates to-be coordinated during class***.

Materials:

Text: Chemistry and Chemical Reactivity (Customized for Loyola University. (2007) by Kotz, Treichel, and Weaver. Please note that the text is a secondary source of information to help clarify concepts presented in lecture. **The primary information is presented in class and also appears on lecture handout materials.**

Calculators will be needed for homework assignments and exams but do not need to be programmable, but should have log/trig functions (typically under $20)

Website: *www.geocities.com/conradnaleway/chem101* **(also found on LUC blackboard)**

Exams:

There will be three 50 minute exams scheduled during the lecture periods and a cumulative final exam. All exams will consist of questions and problems representative of the text and lecture material. All answers to test problems must contain detail information illustrating the steps and method of solution. Answers must contain correct units since this is an essential aspect of the course.

All exams must be signed in the front, upper right hand corner. This signature will be taken as a statement of honest and completely independent work. Instances of academic dishonesty will warrant immediate failure of the course plus referral to the Dean's office.

Exams will be graded and returned as soon as possible, usually the next class period. ALL grading questions, points of clarification and grading errors must be brought to the

instructor's attention during office hours **no later than one week after exam is returned**. There will be no exceptions to this rule! Each returned exam must be copied with original being returned to instructor with a hand written note stapled to exam addressing concern(s).

Exam Grade will be assigned according to the highest percentage computed by the three methods:

a) The average of the three 50 minute class exams, each weighing 1/3, plus completion of the final exam even though no included in grade. **Please note that attendance and completion of the final exam are mandatory and a grade of at least 50% must be achieved!**
b) The average of the top two 50 minute class exams plus the cumulative final. Thus the exams will weigh ¼ each and the final will weigh ½. This relates to dropping the lowest in-class exam.
c) The final exam grade weighed 75% and three exams weighed 25%

NOTE: **Grade is NOT based upon a class curve**. Thus individual performance determines one's grade and is not influenced by other's performance. This thus encourages each student to work collectively to help each other learn. Often discussing and working through a problem with someone else, helps one more than the other person, since it forces one to more critically see through a problem..

Homework Problem Sets: Several sets of problems will be assigned during the semester, roughly one each week. These assignments will largely utilize the OWL homework system which is discussed in detailed in the first few pages of the text (15 %)

Quizzes: Several quizzes (roughly one per discussion period) will be given during the discussion periods (10%)

Final Grade will be based upon:

75% Exam Grade (Above)
10% Discussion Quizzes, and
15% Homework (OWL)

Assignment of Final Grade:

A	**100% - 88%**
B	**87% - 76%**
C	**75% - 60%**
D	**59% - 50%**
F	**<50 %**

TENTATIVE Schedule for Chemistry 101 (1:40pm MWF Fall 2008)

Chapter	Topic	Pages		Class #	Class Date
1	Matter and Measurement	10	57	1,2	8/25, 8/27
2	Atoms and Elements	58	95	3,4	8/29, 9/3
3	Molecules, Ions and Their Compounds	96	139	5,6	9/5, 9/8
4	Chemical Equations and Stoichiometry	140	173	7,8,9	9/10, 9/12,9/15
	EXAM 1			**10**	**17-Sept**
5	Reactions in Aqueous Solutions	174	231	11,12,13	9/19,9/22,9/24
6	Principles of Reactivity	232	293	14.15,16	9/26,9/29,10/1
7	Atomic Structure	294	331	**17,18,19**	10/3,10/8,10/10
	EXAM 2			**20**	**13-Oct**
8	Atomic Electron Configurations	332	371	21-24	10/15 - 10/22
9	Basic Concepts of Chem Bonding	372	435	25-28	10/24 - 10/31
10	Bonding and Molecular Structure: Orbital Hybridization	436	473	29-33	11/31 - 11/12
12	Gases	546	587	34-37	11/14-11/21
	EXAM 3			**38**	**24-Nov**
13	Intermolecular Forces, Liquids & Solids	588	641	39,40	12/1 – 12/3
	REVIEW PERIOD			41	12/5
	FINAL EXAM				**15-Dec**